

Nanocomposites Materials Generated from a Spray.

R.F.C. Marques, J.M.A. Caiut, C.O. Paiva-Santos, S.J.L. Ribeiro, and Y. Messaddeq
Institute of Chemistry, UNESP, Araraquara-SP, 14801-970, Brazil

C. Garcia, D. Neumeyer, H. Dexpert, M. Verelst, and J. Dexpert-Ghys
*Centre d'Elaboration de Matériaux et d'Etudes Structurales UPR 8011,
 CNRS-Université Paul Sabatier, 29 Rue Jeanne Marvig, F 31055 Toulouse Cedex 4, France**
 (Received on 1 July, 2008)

The Spray Pyrolysis (SP) process has been employed at moderate temperature (470-970 K) for the synthesis of micro- and nano-particles of hydrated metal oxides and of related nanocomposites. It has been applied to iron nitrate solutions, without or with the addition of sodium chloride as a flux. After removal of the flux, nanocomposites of ferrihydrite ($5\text{Fe}_2\text{O}_3 \cdot 9\text{H}_2\text{O}$) / hematite ($\alpha\text{-Fe}_2\text{O}_3$) / and an amorphous Fe_2O_3 are obtained, with surface area $140 \text{ m}^2/\text{g}$. It has been applied also to a sol of Aluminum-tri-sec-butoxide, without or with activation by europium or terbium ions. Boehmite ($\gamma\text{-AlOOH}$) powders are synthesized at 470 K, with surface area equal to $180 \text{ m}^2/\text{g}$. Transition alumina ($\gamma\text{-Al}_2\text{O}_3$) are synthesized at 970 K. Nanocomposites made by reaction of the boehmite or $\gamma\text{-Al}_2\text{O}_3$ particles with the amino acid asparagin (ASN) have been prepared and characterized. The luminescence of the Eu^{3+} or Tb^{3+} doped nanocomposites has been investigated: they can be considered as bio-compatible luminescent nanoparticles.

Keywords: Spray-pyrolysis, Nano-powders, Ferrihydrite, Hematite, Boehmite, Alumina.

I. INTRODUCTION

Spray Pyrolysis (SP) is an aerosol process commonly used to form a wide variety of materials in powder form including metals, metal oxides, ceramics, superconductors, etc ... One of the major advantages of SP is the flexibility: pyrolysis may be conducted at high temperatures (1270 K or more) so that high temperature phases may be achieved in one step. Alternatively, the whole process may also be conducted at lower temperature: it may then be described as a rapid drying of the spray. In the past years, we intended to take benefit of this flexibility in order to synthesize sub-micrometric or nanometric particles in a wide range of chemical compositions and considering a wide range of possible applications [1–6] Sub-micronic phosphor powders for plasma display applications were synthesized and post-annealed [1, 2], nanometric powders of zincite ZnO for cosmetic applications were synthesized by SP of mixtures of zinc and lithium (or sodium) nitrates [3]; sub-micronic particles of boehmite (AlOOH) were obtained by spray-drying of an aluminium alcoholate [4], these particles spontaneously dispersed in water to form stable suspensions of boehmite nanoparticles that could be doped with photo luminescent lanthanide ions, then surface-modified to become bio-compatible [5, 6]. The purpose of the present paper is to show how SP, conducted at moderate temperature, may be used for the synthesis of highly divided, highly reactive metal (here Fe or Al) hydroxides and oxides powders.

Iron oxyhydroxides and oxides are of technological importance as catalytic materials, sorbents, pigments, flocculents, coatings, gas sensor and ion exchanger [7]. Among them, ferrihydrite is a natural material present in many near surface environments [8, 9]. Due to its high surface area and reactivity, ferrihydrite plays an important role in the sequestration of contaminants from groundwater. Adsorption on ferrihydrite is commonly used to arsenate sequestration from soils,

groundwater and surface water system [10]. Ferrihydrite is the first phase to precipitate during hydrolysis of Fe^{3+} ions in solution. Its rough formula is $5\text{Fe}_2\text{O}_3 \cdot 9\text{H}_2\text{O}$. Its exact formula is still being debated, as is controversial also its exact crystal structure (see for instance [11, 12]). This is not really surprising since the synthesis conditions have a great impact on the average crystal size and on the reactivity towards water of these nanocrystalline materials. Among the huge literature dealing with structural characterizations, we selected the multiple structural phases model described by Jansen et al. [13] and revisited by Y. Cudennec et al. [14]. Ferrihydrite is metastable and generally considered as goethite $\alpha\text{-FeO}(\text{OH})$ or hematite $\alpha\text{-Fe}_2\text{O}_3$ precursor, depending on pH and/or temperature. Hematite is the end-product of thermal decomposition of ferrihydrite in oxidative atmosphere [15]. Hematite has been demonstrated as a potential catalyst for many environmental application including oxidative removal of carbon monoxide or methane combustion. Currently in the market, NANOCAT superfine Fe_2O_3 nanoparticles (supplied by Mach-I Inc.) are composed of a mixture of amorphous and crystalline structures. The higher catalytic activity of NANOCAT than others commercially available Fe_2O_3 (hematite) is ascribed to its amorphous constituents.

The aluminium oxyhydroxide boehmite $\gamma\text{-AlOOH}$ is the most important precursor of the transition aluminas. Transition aluminas are widely used in the industry of adsorbents and catalysts [16, 17]. $\gamma\text{-AlOOH}$ is traditionally prepared by i) solid state decomposition of gibbsite, ii) precipitation from acidic or basic aluminium aqueous solutions, iii) sol gel procedures from aluminium alcoholates. The third method allows the preparation of very pure boehmite, although from more expensive chemicals. A very common procedure was first described in [18, 19]. It consists of aluminium alkoxide hydrolysis followed by peptization to a clear sol, the gel formation, and then the gel drying to get hydrated boehmite. Further pyrolysis at higher temperature gives porous alumina. The control of the drying conditions is necessary to control the properties of the boehmite powder. The names pseudoboehmite or nano crystalline boehmite are used in con-

*Electronic address: jeannette.dexpert@cemes.fr

trast to microcrystalline or well crystallized boehmite. The nanocrystalline boehmites exhibit crystallite size less than 10 nm, defects and numerous adsorbed water molecules: these features strongly depend on the preparation conditions [20]. The so-called transition aluminas, refer to the group of partially dehydrated aluminium hydroxides other than anhydrous α -alumina (corundum). The diffuse characteristic of the powder patterns reflects a high degree of structural disorder in the transition aluminas. The transition aluminas are widely used in industry as adsorbents, catalysts or catalyst carriers, coatings and soft abrasives because of their fine particle size, high surface area and the catalytic activities of their surfaces. γ - Al_2O_3 is the first transition alumina obtained by thermal decomposition of the boehmite [21, 22].

One aim of this work is to investigate the possibility to obtain by SP α - Fe_2O_3 /ferrihydrite spherical and mesoporous-like structures. Consequently, we propose an alternative material to commercially available NANOCAT product. Nanocomposites particles are generated from a precursor solution containing $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ and NaCl. Herein are investigated the influence of NaCl in the iron nitrate precursor solution and of other operating conditions on the nature, the structure, the morphology and the crystallinity of the powders. Another aim is to synthesize by SP boehmite and transition aluminas as highly divided nanometric powders. The nanometric and/or sub-micronic hydrated aluminas are highly reactive, especially towards amino acids (AA). The nanocomposite hybrids made of boehmite or alumina with the asparagin (ASN) have been studied and characterized as described herein. One possible application of these nanoparticles is luminescent bio-labelling after their activation by an emissive lanthanide ion (Eu^{3+} , Tb^{3+}). Data concerning their optical spectroscopy are described and discussed.

II. EXPERIMENTAL

Samples are prepared by the spray pyrolysis process which is based on an aerosol formation. An intense ultrasonic source close to the liquid surface will produce a geyser, resulting in an aerosol formation due to cavitations effects on the liquid surface. The aerosol, formed by the precursor droplets in an air stream, is conducted through a dry zone at 370-1390 K, then a decomposition-densification zone which can be maintained in the range 570-1470 K. Collection of the dried powders is achieved by a home made electrostatic collector. The ultrasonic generator operates at 2.4 MHz. The adjustable operating parameters are the flow rate (by which the residence time in the furnace is controlled), the temperature in decomposition zone T_{syn} , and of course the nature of the precursor solution. Spheres composed by α - Fe_2O_3 /ferrihydrite nanoparticles are obtained through spray pyrolysis of a precursor solution containing $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ and NaCl as a flux. The idea is to obtain spheres constituted of two phases: the iron oxide phase and a water soluble phase (NaCl). The as-prepared composite spheres are washed several times in water to remove NaCl and the iron oxide is recovered, as schematized in Figure 1. T_{syn} is fixed at 660 or 820 K. For the alumina-based compositions, the precursor solution is prepared following the methodology established by B. Yoldas [18, 19]. A stable sol of Aluminum-tri-sec-butoxide is diluted

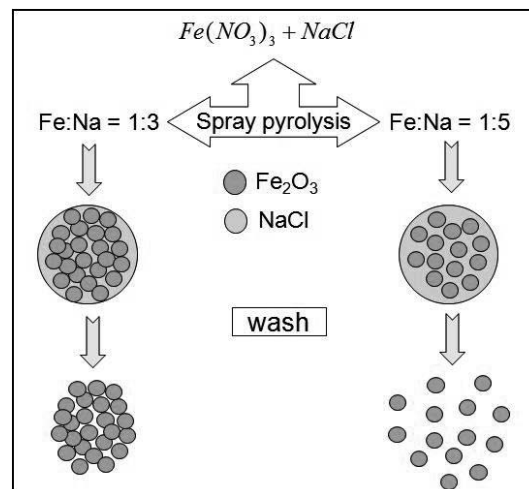


FIG. 1: schematic representation of iron oxide nanoparticles synthesized by SP using a soluble flux (NaCl).

in water to concentration 0.2 mol/L of Al^{3+} , and poured into the vessel over the piezoelectric pellet. T_{syn} is fixed at 470 or 970 K. Activation by optically active ions (Eu^{3+} or Tb^{3+}) is achieved by addition of 2% ($[\text{Eu}]$ or $[\text{Tb}]/[\text{Al}]$) to the distilled water before addition of the aluminum-tri-sec-butoxide. For all SP syntheses, the air flow is 0.3 m^3/h , the residence time in the drying zone is 3-4 seconds and about 6-8 sec in the decomposition-densification zone. Amino-Acids - aluminosilicates nanocomposites have been obtained by dispersing the aluminium oxyhydroxide or oxide particle from SP in asparagines water solutions, under agitation for 12 hours. Adsorption/desorption isotherms have been performed by using a Belsorp-mini (BEL Japan Inc. apparatus), between 0 and 99 P/P₀ of nitrogen and at 77 K after a pre-treatment at 353 K under vacuum during 2 hours. Thermal analyses have been performed in a Labsys TG/DTA (SETARAM) analyser in an alumina crucible at 5 K/min, under oxygen. The size, the morphology and the crystalline structure of the particles have been studied by transmission electron microscopy (TEM) on a Philips CM12 microscope. The crystalline structures of samples have been investigated by X ray powder diffraction (Seifert XRD 3000 diffractometer) and the quantitative phases analysis obtained by Rietveld refinements [23] using the program GSAS [24]. In order to obtain an internal standard without microstrain and with a high average crystallite size, the 99.99 % cerium oxide is fired at 1570 K for 48 h. Room temperature luminescence excitation and emission spectra have been measured by a Hitachi F1100 spectrofluorimeter at 2.5 nm spectral resolution.

III. RESULTS AND DISCUSSION

A. Iron oxyhydroxides and oxides.

Sample FE1 has been prepared without NaCl in the precursor solution. As expected in spray pyrolysis synthesis under the selected operating conditions, the particles are spherical and their diameters fall in the range 100 to 2000 nm. One TEM image of sample FE1 is displayed in figure 2a. All

sample	[Fe]	[Na]	Fe:Na	T _{syn}	XRD
FE1	0.15	—	—	820	HEM
FE2	1.0	—	—	820	HEM
FE3	1.5	—	—	820	HEM
FENA1	0.15	0.45	1:3	820	HEM+FEH+NaCl
FENA2	0.15	0.45	1:3	660	HEM+FEH+NaCl+NaNO ₃
FENA3	0.10	0.50	1:5	820	HEM+FEH+NaCl

TABLE I: experimental synthesis conditions for iron-based samples: [Fe] and [Na] in mol.L⁻¹, T_{syn} in K, XRD: crystalline phases identified (HEM= α -Fe₂O₃, FEH=ferrihydrite)

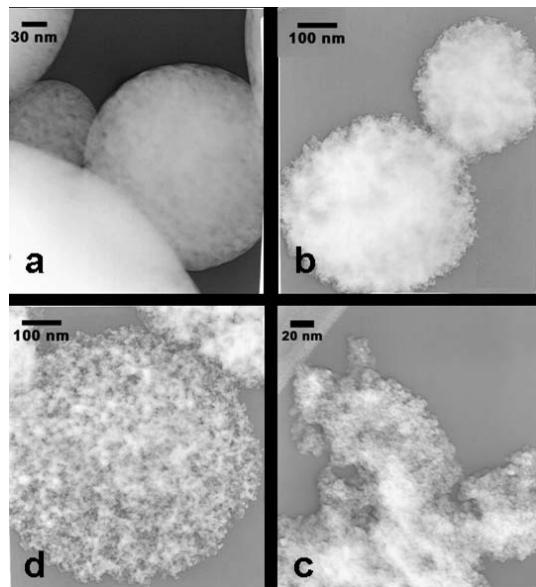


FIG. 2: TEM images of samples FE1(a); FeNa2(b); FENA2-washed (c), and FENA1-washed (d)

XRD peaks observed for this sample may be indexed with hematite (α -Fe₂O₃- PDF no 89-2810), as shown in figure 3. The XRD patterns of as obtained samples FENA 1-3 display well defined narrow peaks superposed on broad components. Hematite is detected, together with NaCl for the three samples and with NaNO₃ in sample FENA2 synthesized at 660 K. This results from the reaction of nitrate species freed by the iron precursor with sodium ions. Nitrate species are evolved when the synthesis is conducted at higher temperature (820 K). These observations prove the complexity of the SP synthesis during which occur in a limited range of time the solute precipitation, the solvent evaporation and chemical reactions between the constituents in each droplet [25]. After water washing, XRD patterns of soluble phases NaCl and NaNO₃ are not observed, the diffraction diagrams exhibit the peaks of α -Fe₂O₃ and broad features assigned to hydrated species (ferrihydrites) as discussed below.

The particles of samples FENA1-3 are spherical and sub-micrometric. The spheres observed are nanocomposites made of the iron salts and the sodium salts (Figure 2b). For sample FENA3, obtained with the Fe:Na ratio 1:5, some NaCl cubic-like aggregates are also observed. These cubic aggregates are no more observed after washing the sample in water. After wash, highly porous aggregates, retaining the spherical shape of the nanocomposite particles are observed by TEM for FENA1 (Figure 2d), but hardly seen for FENA2,

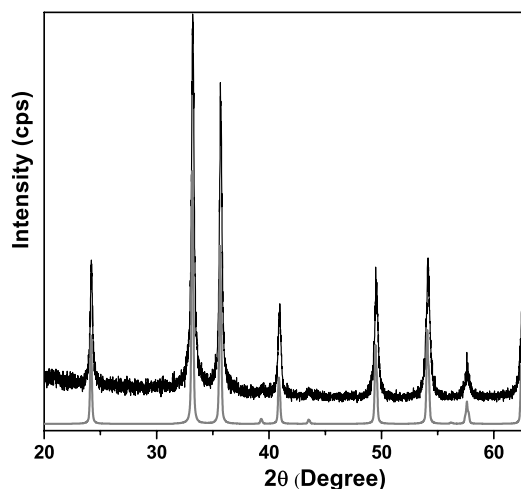


FIG. 3: XRD analysis. Sample FE1 and calculated diagram for hematite α -Fe₂O₃.

washed (Figure 2c), same for FENA3 washed (not shown). These observations suggest first that at higher T_{syn}, (820 for SPNA1 versus 660 K for SPNA2), the individual nanoparticles start to coalesce during the spray pyrolysis, and the formation of necks prevents desagglomeration of the spheres during washing. Another important parameter is the Fe:Na ratio in the precursor: despite of being obtained at T_{syn}= 820 K, aggregates of nanoparticles in FENA3 after wash are no more spherical, suggesting that the high NaCl concentration has not allowed enough contact among iron-based nanoparticles so that necks could not be formed prior to dissolution of the NaCl flux. Therefore, by carefully adjusting Fe:Na ratio and furnace temperature, sub-micronic spheres of iron oxides nanoparticles can be obtained through spray pyrolysis. The N₂ adsorption-desorption isotherms of samples FENA1 (washed) have been measured: it comes that the surface area is 140 ± 15 m²/g. Compared to sample FE1, the XRD pattern recorded for sample FENA1 after wash exhibits significantly broadened peaks which can be assigned to poorly crystalline ferrihydrite, superimposed to the narrower diffraction peaks of hematite. Hematite is the end-product of thermal decomposition of the complex ferrihydrite. At same T_{syn} (820 K), the reaction is then less completed following the flux-assisted SP process than following the more conventional one phase SP. A more precise and quantitative description of the phases present in the SPNA1, after wash, has been achieved. A precisely weighted standard (CeO₂ at 11.4 % w/w) has been added to the sample in order to normalize the different fractions calculated by Rietveld analysis routine. In this model (from references [13, 14]), the structure of ferrihydrite is considered as being a superposition of two components: a defect-free phase of trigonal space group (P-31c) and a defective phase with a trigonal P3 space group. Rietveld analysis results are presented in Table II and in Figure 4: besides 27 weight % of hematite and 32 weight % of ferrihydrite, it is concluded that 41 weight % of the sample are amorphous. Taking the general formula 5Fe₂O₃·9H₂O for the ferrihydrite, it comes from the Rietveld quantitative analysis that 5.4 weight % of the sample FENA1(washed) is H₂O. Experimentally, the total weight loss of the sample measured by

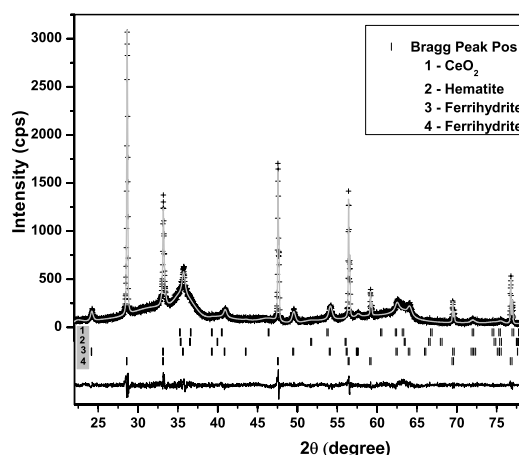


FIG. 4: XRD analysis. Sample FENA1 washed + internal standard CeO_2 ; Rietveld simulation.

thermal analysis between 300 and 920 K is 7 ± 0.14 weight %, which is consistent with the previous estimation. In turn, it proves that the composition of the amorphous phase is anhydrous Fe_2O_3 .

phase	amorphous	HEM	FEH(P3)	FEH (P-31c)
weight%	41	27	11	21
H_2O weight%	—	—	1.85	3.55

TABLE II: sample FENA1 (washed): Refined phases weight% HEM = α - Fe_2O_3 , FEH=ferrihydrite

B. Aluminium oxyhydroxides and oxides.

The synthesis method selected to elaborate by SP the aluminium based samples is in fact the fast drying of a colloidal solution of boehmite, in order to stop the sol-gel process prior to the long range ordering of the solid phase. The elaboration conditions, the crystal phases identified and some typical properties of the samples are gathered in table III. As for the previous case, the SP synthesis produces NPs agglomerated in sub-micronic spheres. The measured surface area is high: $180 \pm 15 \text{ m}^2/\text{g}$ (sample AL200). The sample elaborated at the lower temperature (AL200) exhibits the XRD pattern of nanocrystalline boehmite γ - AlOOH (Figure 5). It may be spontaneously dispersed in water. In ref.[5] we have shown that the rapid drying (a few seconds) of the sol stops the development of the aluminium hydroxide network at the level of a few crystal cells. Half the hydroxyl groups are in fact at the external surface of the crystallites. This resulted in a very large number of surface hydroxyls free for water adsorption, actually the chemical composition of SP nano crystallized boehmite is $\text{AlOOH} \cdot (1.18 \pm 0.24) \text{H}_2\text{O}$. The very hydrophilic nature of the surface of SP-boehmite crystallites is at the origin of their ability to form stable suspensions in water. When the synthesis is performed at higher temperature (AL700), the XRD peaks are still broader and fit well with what is reported in literature as the transition alumina γ - Al_2O_3 . AL700 powder precipitates in water. Nanocomposites obtained by reac-

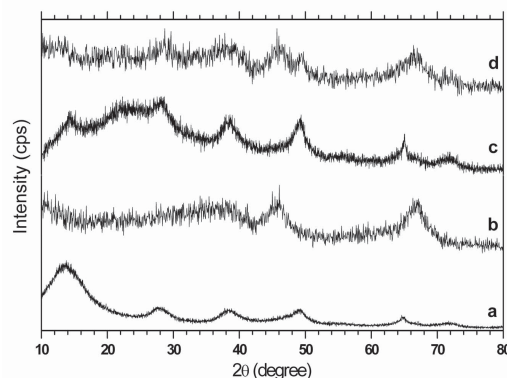


FIG. 5: XRD patterns of samples AL200 (a), AL700 (b), AL200-ASN (c), AL700-ASN (d).

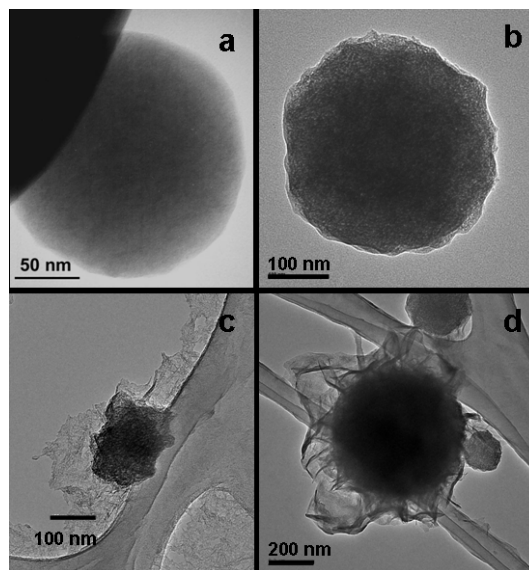


FIG. 6: TEM images of samples AL200 (a), AL700 (b), AL200-ASN (c), AL700-ASN (d).

tion with ASN behave the same way with respect to dispersion. The interactions between the ASN and the particles surfaces are imaged in Figure 6. The AL200 (boehmite) spheres have a smooth surface (Figure 6a), whereas the AL700 (γ - Al_2O_3) are less regular, and denser (Figure 6b). After reaction with ASN at molar ratio $\text{ASN}:\text{Al} = 0.3:1$ and drying, most of the boehmite spheres have disaggregated and the sample looks like a film in which are dispersed nanoparticles. Aluminium is identified by local analysis (EDXS in TEM) for all parts of the film, and the XRD is that of boehmite. The TEM image selected here (Figure 6c) shows the composite film and one mineral bigger particle not completely destroyed. For the transition aluminas AL700-ASN (Figure 6d), the mixing of the ASN and of the mineral part is less intimate: one observes an organic film wrapping the sub micronic γ - Al_2O_3 particles.

The peculiar emission of Eu^{3+} in AL200:Eu (atomic ratio $\text{Eu}:\text{Al} = 0.02:1$) is very well preserved in the composite AL200:Eu-ASN ($\text{Eu}:\text{Al} = 0.02:1$, $\text{ASN}:\text{Al} = 0.3:1$) displayed in Figure 7a. In the same way, the emission spectra of Eu^{3+} in AL700:Eu and in AL700:Eu-ASN samples are very similar (the last one displayed in Figure 7b). One recognizes easily

- A.Garcia, P. Guillot, J. Galy, R. Mauricot and S. Schamm, J. Phys. D, **38**, 3261 (2005).
- [3] C. Rossignol, M. Verelst, J. Dexpert-Ghys and S. Rul, Adv. Sci. Tech. **45**, 237 (2006).
- [4] J.M.A. Caiut, J. Dexpert-Ghys, Y. Kihn, M. Verelst, H. Dexpert, S.J.L. Ribeiro and Y. Messaddeq, Powder Techn. **190**, 95-98 (2009).
- [5] J.M.A. Caiut, S.J.L. Ribeiro, Y. Messaddeq, J. Dexpert-Ghys, M. Verelst and H. Dexpert, Nanotech. **18**, 455605 (2007).
- [6] J.M.A. Caiut, L. Bazin, R. Mauricot, H. Dexpert, S.J.L. Ribeiro and J. Dexpert-Ghys, J. Non Cryst. Sol. **35**, 4860-4864 (2008).
- [7] F. Schuth, Chem. Mater. **13**, 3184 (2001).
- [8] D. G. Rancourt, D. Fortin, T. Pichler, P.J. Thibault, G. Lamarche, R.V. Morris and P.H.J. Mercier, Am. Mineral. **86**, 834 (2001).
- [9] U. Schwertmann, L. Carlson and E. Murad, Clays Clay Miner. **35**, 297 (1987).
- [10] Y. Jia, L. Xu, X. Wang and G.P. Demopoulos, Geochim. et Cosmochim. Acta **71**, 1643 (2007).
- [11] R.L. Penn, Science **316**, 1704 (2007).
- [12] F.M. Michel, L. Ehm, S.M. Antao, P.L. Lee, P.J. Chupas, G. Liu, D.R. Strongin, M.A. A. Schoonen, B.L. Phillips and J.B. Parise, Science **316**, 1726 (2007).
- [13] E. Jansen, A. Kyek, W. Schfer, U. Schwertmann, Appl. Phys. A, **74**, 1004 (2002).
- [14] Y. Cudennec, A. Lecerf, J. Solid State Chem. **179**, 716 (2006).
- [15] R.M. Cornell and U. Schwertmann, The Iron Oxides, Structure, Properties, Reactions, Occurrence and Uses (WCH, Weinheim, New York, Basel, Cambridge, Tokyo, (2004)).
- [16] A. Pearson, Aluminium oxide, activated, in Kirk-Othmer Encyclopedia of Chemical Technology (J. Wiley, New York, 2, 291 (1994)).
- [17] C. Misra, Aluminium oxide, hydrated, in Kirk-Othmer Encyclopedia of Chemical Technology (J. Wiley, New York, 2, 317 (1994)).
- [18] B. E. Yoldas, Ceramic Bull. **54:3**, 289 (1975).
- [19] B. E. Yoldas, J. Mater. Sci. **10**, 1856 (1975).
- [20] M. Nguefack, S. Rossignol, A.F. Popa, and C. Kappenstein, Phys. Chem. Chem. Phy. **5**, 4279 (2003).
- [21] R.S. Zhou, R.L. Snyder, Appl. Crys. **47**, 617 (1991).
- [22] K. & MISRA, C.. Alcoa Technical Paper 19, revised. Alcoa Laboratories, Pittsburg, PA, USA. (1987)
- [23] H.M. Rietveld, J. Appl. Crystall. **2**, 65 (1969).
- [24] A.C. Larson and R.B. Von Dreele, Los Alamos National Laboratory. Los Alamos , EUA. Copyright, 1985-2000, The Regents of the University of California, 2001.
- [25] N. Reuge, J. Dexpert-Ghys, M. Verelst, H. Dexpert and B. Caussat, AIChE J. **54**, 394 (2008).